

Formula and Equation Balancing

Exercise #1

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Follow the steps! →

1. Write subscripts below any free diatomic elements.
2. Balance the formulas using oxidation numbers and subscripts.
3. Balance the equation using coefficients.

Name _____ Date _____ Class Period _____

UNBALANCED

1.	$\text{Na} + \text{O} \rightarrow \text{NaO}$
2.	$\text{Na} + \text{Cl} \rightarrow \text{NaCl}$
3.	$\text{H} + \text{O} \rightarrow \text{HO}$
4.	$\text{Mg} + \text{O} \rightarrow \text{MgO}$
5.	$\text{Al} + \text{O} \rightarrow \text{AlO}$
6.	$\text{Ca} + \text{F} \rightarrow \text{CaF}$
7.	$\text{P} + \text{O} \rightarrow \text{PO}$
8.	$\text{Li} + \text{H}_2\text{O} \rightarrow \text{LiO} + \text{H}$
9.	$\text{N} + \text{H} \rightarrow \text{NH}$
10.	$\text{KClO}_3 \rightarrow \text{KCl} + \text{O}$
11.	$\text{Cu} + \text{AgNO}_3 \rightarrow \text{CuNO}_3 + \text{Ag}$
12.	$\text{BaCl} + \text{HSO}_4 \rightarrow \text{BaSO}_4 + \text{HCl}$
13.	$\text{NaS} + \text{AgNO}_3 \rightarrow \text{NaNO}_3 + \text{AgS}$
14.	$\text{Fe} + \text{O} \rightarrow \text{FeO}$
15.	$\text{Zn} + \text{HCl} \rightarrow \text{ZnCl} + \text{H}$

BALANCED
